

Name: _____

Date: _____

Period: _____

Chemistry Unit 3 Review

Part 1: History of the Atom

Scientist/Model name	Drawing/Illustration	Could account for	Could NOT account
Dalton		<i>All matter is made of atoms. Atoms are small and indestructible. All atoms of a given element are identical in mass and properties</i>	
JJ Thomson		<i>Discovered electrons</i>	<i>Stability of atom</i>
Rutherford		<i>Discovered proton</i>	<i>Stability of atom</i>
Bohr		<i>Energy levels</i>	<i>Continuous and discontinuous spectrum</i>

Answer the following questions:

Time Line review

Who was the ancient Greek philosopher who first proposed the notion of the atom? *Democritus*

Who's model first introduced the concept of energy levels? *Bohr*

1. What were the major problems of Dalton's atomic theory? *His first two postulates were incorrect.*
2. Whose model was nicknamed the plum pudding model? *Thomson*
3. Which experiment verified that atoms could NOT be solid and the center had a positive charge? *Rutherford*
4. What did Rutherford's gold foil experiment show about the atom? *The nucleus was positively charged.*
5. Rutherford's model of the atom was most similar to Bohr model.
6. What atomic models did not have empty space? *Thomson*

Chemistry Unit 3 Review Answers

Xiaolong Qi

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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