

Thermochemistry Quiz 1

Questions 1-4 (1 pt. each); Questions 5-6 (2 pts. Each)

- _____ 1. An endothermic reaction causes the surroundings to
A. warm up.
B. become acidic.
C. condense.
D. decrease in temperature.
E. release CO₂
- _____ 2. According to the first law of thermodynamics:
A. energy is neither lost nor gained in any energy transformations.
B. perpetual motion is possible.
C. energy is conserved in quality but not in quantity.
D. energy is being created as time passes. We have more energy in this universe now than when time began.
- _____ 3. A calorimeter
A. is equal to the molar enthalpy of a reaction.
B. is a dieting aid.
C. is an indicator of a spontaneous reaction.
D. is a device used to measure the transfer of heat energy.
E. is useful in measuring the amount of heat released by endothermic reactions.
- _____ 4. The specific heat of gold is 0.129 J/g•°C. What is the molar heat capacity of gold?
A. 0.039 J/mol•°C
B. 0.129 J/mol•°C
C. 25.4 J/mol•°C
D. 39.0 kJ/mol•°C
E. 197 J/mol•°C
- _____ 5. Copper metal has a specific heat of 0.385 J/g•°C. Calculate the amount of heat required to raise the temperature of 22.8 g of Cu from 20.0°C to 875°C.
A. 1.97×10^{-2} J
B. 1.0×10^{-2} J
C. 329 J
D. 7.51 kJ
E. 10.5 kJ
- _____ 6. The combustion of butane produces according to the equation
$$2\text{C}_4\text{H}_{10}(\text{g}) + 13\text{O}_2(\text{g}) \rightarrow 8\text{CO}_2(\text{g}) + 10\text{H}_2\text{O}(\text{l}) \quad \Delta H^\circ_{\text{rxn}} = -5,314 \text{ kJ}$$

How many grams of butane must be burned to release 6,375 kJ of heat?
A. 1.20 g
B. 139 g
C. 0.0413 g
D. 69.7 g
E. 97.8 g

Questions 7-8: Label the following processes as either *exothermic* or *endothermic*. (1 pt. each)

7. Dry ice (solid CO₂) sits at room temperature and sublimates to CO₂ gas. _____
8. A block of solid aluminum is heated until it melts. _____

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Table of Contents Chemistry Thermochemistry Quiz Questions

1. Understanding the eBook Chemistry Thermochemistry Quiz Questions
 - The Rise of Digital Reading Chemistry Thermochemistry Quiz Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Thermochemistry Quiz Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Thermochemistry Quiz Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Thermochemistry Quiz Questions
 - Personalized Recommendations
 - Chemistry Thermochemistry Quiz Questions User Reviews and Ratings
 - Chemistry Thermochemistry Quiz Questions and Bestseller Lists
5. Accessing Chemistry Thermochemistry Quiz Questions Free and Paid eBooks
 - Chemistry Thermochemistry Quiz Questions Public Domain eBooks
 - Chemistry Thermochemistry Quiz Questions eBook Subscription Services
 - Chemistry Thermochemistry Quiz Questions Budget-Friendly Options
6. Navigating Chemistry Thermochemistry Quiz Questions eBook Formats

- ePub, PDF, MOBI, and More
 - Chemistry Thermochemistry Quiz Questions Compatibility with Devices
 - Chemistry Thermochemistry Quiz Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Thermochemistry Quiz Questions
 - Highlighting and Note-Taking Chemistry Thermochemistry Quiz Questions
 - Interactive Elements Chemistry Thermochemistry Quiz Questions
 8. Staying Engaged with Chemistry Thermochemistry Quiz Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Thermochemistry Quiz Questions
 9. Balancing eBooks and Physical Books Chemistry Thermochemistry Quiz Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Thermochemistry Quiz Questions
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Thermochemistry Quiz Questions
 - Setting Reading Goals Chemistry Thermochemistry Quiz Questions
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Thermochemistry Quiz Questions
 - Fact-Checking eBook Content of Chemistry Thermochemistry Quiz Questions
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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