

## Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H<sub>2</sub>O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

|    |              |
|----|--------------|
| Na | 23.0         |
| Cl | 35.5         |
|    | <hr/> 58.5 g |

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

|    |            |
|----|------------|
| Na | 23.0       |
| H  | 1.0        |
| C  | 12.0       |
| O  | 48.0       |
|    | <hr/> 84.0 |

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

|    |             |
|----|-------------|
| Fe | 111.6       |
| O  | 48.0        |
|    | <hr/> 159.6 |

# Chemistry U5 Ws1 V3 0 Answer Key

**E. N. Ramsden**



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## **Table of Contents Chemistry U5 Ws1 V3 0 Answer Key**

1. Understanding the eBook Chemistry U5 Ws1 V3 0 Answer Key
  - The Rise of Digital Reading Chemistry U5 Ws1 V3 0 Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry U5 Ws1 V3 0 Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry U5 Ws1 V3 0 Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry U5 Ws1 V3 0 Answer Key
  - Personalized Recommendations
  - Chemistry U5 Ws1 V3 0 Answer Key User Reviews and Ratings
  - Chemistry U5 Ws1 V3 0 Answer Key and Bestseller Lists
5. Accessing Chemistry U5 Ws1 V3 0 Answer Key Free and Paid eBooks

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- Chemistry U5 Ws1 V3 0 Answer Key eBook Subscription Services
- Chemistry U5 Ws1 V3 0 Answer Key Budget-Friendly Options
- 6. Navigating Chemistry U5 Ws1 V3 0 Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry U5 Ws1 V3 0 Answer Key Compatibility with Devices
  - Chemistry U5 Ws1 V3 0 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry U5 Ws1 V3 0 Answer Key
  - Highlighting and Note-Taking Chemistry U5 Ws1 V3 0 Answer Key
  - Interactive Elements Chemistry U5 Ws1 V3 0 Answer Key
- 8. Staying Engaged with Chemistry U5 Ws1 V3 0 Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry U5 Ws1 V3 0 Answer Key
- 9. Balancing eBooks and Physical Books Chemistry U5 Ws1 V3 0 Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry U5 Ws1 V3 0 Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry U5 Ws1 V3 0 Answer Key
  - Setting Reading Goals Chemistry U5 Ws1 V3 0 Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry U5 Ws1 V3 0 Answer Key
  - Fact-Checking eBook Content of Chemistry U5 Ws1 V3 0 Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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