

UNIT 6: CHEMICAL REACTIONS TEST REVIEW

Name: _____

Date: _____ Hour: _____

Part 1: General Knowledge Questions: Answer the following questions regarding information learned from unit 6.

- Fill in the blanks with what categories of substances go where in a chemical reaction:
_____ $\rightarrow$ _____
- Give five lines evidence that a chemical reaction has occurred: _____,
_____, _____ & _____.
- Identify the seven diatomic molecules: _____, _____, _____, _____, _____, _____, & _____.
- In the expression, $5 \text{H}_2\text{O}$, the number five represents a _____ and the number 2 represents a _____. (coefficient or subscript for each blank)
- According to the law of conservation of mass, in a chemical reaction...(select one)
 - the types of atoms on either side of the equation must be different.
 - the number of each type of atom on either side of the equation must be equal.
 - the bonds in chemical compounds cannot be broken.
 - a compound that is found in the reactants must also be found in the products.
- What type of chemical reaction involves the cation of one compound switching places with a cation of a second compound? _____
- What type of chemical reaction has one reactant and multiple products? _____
- What type of chemical reaction involves a hydrocarbon and oxygen gas as reactants? _____
- What type of chemical reaction has multiple reactants but a single product? _____
- What type of chemical reaction involves a free element switching places with one ion in a compound?

- Identify the type of chemical reactions shown below:
 - _____ $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$
 - _____ $\text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{O} + \text{CO}_2$
 - _____ $2 \text{H}_2 + \text{O}_2 \rightarrow 2 \text{H}_2\text{O}$
 - _____ $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
 - _____ $\text{Zn} + 2 \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
 - _____ $\text{HBr} + \text{NaOH} \rightarrow \text{NaBr} + \text{HOH}$
 - _____ iron (III) hydroxide reacts with nitric acid
 - _____ calcium hydroxide (heated)
 - _____ propane(C_3H_8) burns in oxygen
 - _____ lead (II) bromide reacts with zinc
 - _____ calcium reacts with fluorine
- In the reaction: $\text{H}_2\text{SO}_4 + 2 \text{NaOH} \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$, what is the ratio of sodium hydroxide to sodium sulfate? _____

Chemistry Unit 6 Test Chemical Reactions

IM Harris



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Table of Contents Chemistry Unit 6 Test Chemical Reactions

1. Understanding the eBook Chemistry Unit 6 Test Chemical Reactions
 - The Rise of Digital Reading Chemistry Unit 6 Test Chemical Reactions
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Unit 6 Test Chemical Reactions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Unit 6 Test Chemical Reactions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Unit 6 Test Chemical Reactions
 - Personalized Recommendations
 - Chemistry Unit 6 Test Chemical Reactions User Reviews and Ratings
 - Chemistry Unit 6 Test Chemical Reactions and Bestseller Lists

5. Accessing Chemistry Unit 6 Test Chemical Reactions Free and Paid eBooks
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 - Chemistry Unit 6 Test Chemical Reactions eBook Subscription Services
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6. Navigating Chemistry Unit 6 Test Chemical Reactions eBook Formats
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 - Chemistry Unit 6 Test Chemical Reactions Compatibility with Devices
 - Chemistry Unit 6 Test Chemical Reactions Enhanced eBook Features
7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Chemistry Unit 6 Test Chemical Reactions
 - Interactive Elements Chemistry Unit 6 Test Chemical Reactions
8. Staying Engaged with Chemistry Unit 6 Test Chemical Reactions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Unit 6 Test Chemical Reactions
9. Balancing eBooks and Physical Books Chemistry Unit 6 Test Chemical Reactions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Unit 6 Test Chemical Reactions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Unit 6 Test Chemical Reactions
 - Setting Reading Goals Chemistry Unit 6 Test Chemical Reactions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Unit 6 Test Chemical Reactions
 - Fact-Checking eBook Content of Chemistry Unit 6 Test Chemical Reactions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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