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B 7

Multiple Choice: Circle the most correct answer. (8K, 7T)

1. How many pi (π) bonds are present in the molecule depicted below?

- A) 2 B) 3 C) 4 D) 5 E) 1

2. Double bonds can be thought of as consisting of:

- A) Two sigma (
- σ
-) bonds B) One pi (
- π
-) bond
-
- C) A sigma (
- σ
-) bond and a pi (
- π
-) bond D) One sigma (
- σ
-) bond
-
- E) Two pi (
- π
-) bonds

3. Covalent bonds are most commonly found between:

- A) elements with similar but relatively low first ionization energies.
-
- B) elements with low first ionization energies.
-
- C) elements with a large difference in first ionization energies.
-
- D) high atomic mass metals and low atomic mass non-metals.
-
- E) elements with relatively high first ionization energies.

4. A difference between ionic and molecular compounds is that ionic compounds:

- A) dissolve in water, but molecular compounds are insoluble.
-
- B) usually form crystals, whereas molecular compounds usually do not.
-
- C) conduct electricity in the solid state whereas molecular compounds usually do not.
-
- D) usually melt at higher temperatures than molecular compounds.

5. Which one of the following is UNLIKELY to be true of an ionic compound?

- A) It exists as a crystalline compound at room temperature.
-
- B) In the pure liquid state, it is a poor conductor of electricity.
-
- C) It is a good conductor of electricity in aqueous solutions.
-
- D) It has a melting point above
- 250°C
- .
-
- E) It is a compound of a metal and a non-metal.

6. What is the shape of the nitrate ion, NO_3^{-} ?

- A) linear B) trigonal planar C) tetrahedral D) octahedral E) trigonal pyramidal

Covalent Bonding Quiz Answers

RD Boyd



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