

Name: _____

Naming Practice

When dealing with compounds, the first thing that should be done is determine if they are ionic or covalent. And if I am having you name them and it has a metal, then it is ionic. If you see nonmetals then it is covalent.

Dealing with Ionic Compounds

Since ionic compounds have ions, these are the ones that you have to worry about charges with. Most of the time you can use the periodic table to determine the charge. If the element is from the first column of the periodic table, it has a +1 charge. Second column = +2 charge. Group 13 (Boron Family) = +3 charge. Group 15 = -3 charge. Group 16 = -2 charge. Group 17 = -1 charge. If it is a transition element, inner transition, or Group 14, we will assume they have more than one charge, and we will say what charge it is by using a roman numeral.

Ionic Compounds also deal more with the polyatomic ions. If you see -ate or -ite in the name, you are dealing with a polyatomic. If you are dealing with an element, you will use the ending -ide. The tricky polyatomic ions are ammonium (NH_4^+), cyanide (CN^-), hydroxide (OH^-), and peroxide (O_2^{2-}) (Their formulas go in the blanks). The final thing you have to worry about is switchy-switchy, which is switching the charge of one ion to be the subscript of the other part of the compound, and reduce the numbers if possible. Remember, you only take the number and not the sign. Let's practice:

Compound Name	Ions that make up the compound	Resulting Formula
calcium chloride	$Ca^{+2} Cl^{-1}$	$CaCl_2$
cobalt (V) nitride	$Co^{+5} N^{-3}$	Co_3N_5
strontium hydroxide	$Sr^{+2} (OH)^{-1}$	$Sr(OH)_2$
lead (III) sulfate	$Pb^{+3} (SO_4)^{-2}$	$Pb_2(SO_4)_3$
iron (III) phosphate	$Fe^{+3} (PO_4)^{-3}$	$FePO_4$ (reduces)
magnesium fluoride	$Mg^{+2} F^{-1}$	MgF_2
chromium (III) sulfite	$Cr^{+3} (SO_3)^{-2}$	$Cr_2(SO_3)_3$
barium selenide	$Ba^{+2} Se^{-2}$	$BaSe$
osmium (IV) oxide	$Os^{+4} O^{-2}$	OsO_2

reverse switchy-switchy doesn't work on these two.
start w/ anion 1st (from P.T.)

Covalent Bonding Packet Answers

Sabine Zange



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Table of Contents Covalent Bonding Packet Answers

1. Understanding the eBook Covalent Bonding Packet Answers
 - The Rise of Digital Reading Covalent Bonding Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonding Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Bonding Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonding Packet Answers
 - Personalized Recommendations

- Covalent Bonding Packet Answers User Reviews and Ratings
- Covalent Bonding Packet Answers and Bestseller Lists
- 5. Accessing Covalent Bonding Packet Answers Free and Paid eBooks
 - Covalent Bonding Packet Answers Public Domain eBooks
 - Covalent Bonding Packet Answers eBook Subscription Services
 - Covalent Bonding Packet Answers Budget-Friendly Options
- 6. Navigating Covalent Bonding Packet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Covalent Bonding Packet Answers Compatibility with Devices
 - Covalent Bonding Packet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Covalent Bonding Packet Answers
 - Highlighting and Note-Taking Covalent Bonding Packet Answers
 - Interactive Elements Covalent Bonding Packet Answers
- 8. Staying Engaged with Covalent Bonding Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Covalent Bonding Packet Answers
- 9. Balancing eBooks and Physical Books Covalent Bonding Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Covalent Bonding Packet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Covalent Bonding Packet Answers
 - Setting Reading Goals Covalent Bonding Packet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Covalent Bonding Packet Answers
 - Fact-Checking eBook Content of Covalent Bonding Packet Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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