

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization constant | 5. K_a | 6. K_b |
| | 7. K_w | |

Key Equation

- | | |
|-------------------------------------|---|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}_3\text{O}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}_3\text{O}^+]$ |

Part A Completion

Write the requested unit or numerical information of the values provided that are indicated in parentheses. Significant figures are indicated with the appropriate symbol.

1. What substance is _____, a weak hydrogen ion donor?
2. _____ is the concentration of H^+ ions present in 0.00100 mol/L of _____.
3. _____ is the concentration of _____ in 0.00100 mol/L of _____.
4. _____ is the concentration of _____ in 0.00100 mol/L of _____.
5. _____ is the concentration of _____ in 0.00100 mol/L of _____.
6. _____ is the concentration of _____ in 0.00100 mol/L of _____.
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14. _____ is the concentration of _____ in 0.00100 mol/L of _____.
15. _____ is the concentration of _____ in 0.00100 mol/L of _____.
16. _____ is the concentration of _____ in 0.00100 mol/L of _____.
17. _____ is the concentration of _____ in 0.00100 mol/L of _____.
18. _____ is the concentration of _____ in 0.00100 mol/L of _____.
19. _____ is the concentration of _____ in 0.00100 mol/L of _____.
20. _____ is the concentration of _____ in 0.00100 mol/L of _____.

1. _____
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