

REACTION CATEGORY	DOUBLE REPLACEMENT
REACTION DESCRIPTION	During double replacement, the cations and anions of two different compounds switch places.
REACTION FORMAT	$AB + CD \rightarrow AD + CB$
REACTION GUIDELINES	1. It is important that the formulas of the products be written correctly. If they are correct, balancing the equation is a simple task; if not, the equation will probably never balance. 2. In these reactions, there is never a change in oxidation state. 3. Sometimes you must determine if a reaction actually takes place? For example : Does a mixture of NaCl and H₂SO₄ react to give Na₂SO₄ and HCl, or rather, does a mixture of Na₂SO₄ and HCl react to give NaCl and H₂SO₄. Obviously we cannot test every reaction before we write the equation, but fortunately, there are certain conditions under which a reaction goes to completion (i.e. goes in one direction only). These are summarized below. A reaction takes place or tends to go to completion if: - One of the products is a gas and is allowed to escape. - An unionized substance such as H₂O or NH₃ is formed. - An insoluble substance is formed. The first two of these are obvious if we are able to recognize which substances are gases. The most common inorganic gases are H₂, Cl₂, O₂, N₂, H₂S, HF, HCl, HBr, HI, CO, CO₂, SO₂, SO₃, NH₃, NO, N₂O, NO₂ and HCN. The most difficult aspect of reactions of this type is the ability to recognize insoluble substances. Here are some solubility guidelines: - All nitrates and acetates are soluble. - All chlorides, bromides, and iodides are soluble except those of Pb²⁺, Ag⁺, and Hg²⁺. - All sulfates are soluble except those of Ba²⁺, Sr²⁺, and Pb²⁺. CaSO₄, Ag₂SO₄, and Hg₂SO₄ are slightly soluble. - All hydroxides are insoluble except those of group I in the periodic table, NH₄⁺, and Ba²⁺. Ca(OH)₂ and Sr(OH)₂ are slightly soluble. - All carbonates and phosphates are insoluble except those of group I and NH₄⁺. Many hydrogen phosphates are soluble. - All sulfides are insoluble except those of Group I and Group II in the periodic table and NH₄⁺. In addition, keep in mind the following when H₂CO₃, H₂SO₃, and NH₄OH are formed as products: - H₂CO₃ decomposes into CO₂ + H₂O - H₂SO₃ decomposes into SO₂ + H₂O - NH₄OH decomposes into NH₃ + H₂O
REACTION GUIDELINE EXAMPLES	$AgNO_3 + NaCl \rightarrow AgCl + NaNO_3$ $CaCO_3 + HCl \rightarrow CaCl_2 + CO_2 + H_2O$ (#7) $Pb(NO_3)_2 + CuSO_4 \rightarrow PbSO_4 + Cu(NO_3)_2$

Double Replacement Reaction 9 2 Worksheet

Ian Pickup



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