

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa, from K_w .
2. Calculate the pH of a solution of a weak acid or base.
3. Calculate the pH of a solution of a salt.

Vocabulary

- acid dissociation constant
- base dissociation constant

- conjugate acid–base pair
- hydrolysis
- indicator

- neutralization
- polyprotic acid

Key Equations

$$K_a = \frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]}$$

(see Table 19.2)

$$K_b = \frac{[\text{B}][\text{OH}^-]}{[\text{BOH}]}$$

(see Table 19.2)

Part A: Completion

Use the appropriate words to complete the following statements. Write the word or phrase in the blank space provided. Do not use the same word or phrase more than once.

1. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
2. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is _____.
3. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
4. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is _____.
5. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
6. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is _____.
7. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
8. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is _____.

9. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
10. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is _____.
11. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
12. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is _____.
13. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
14. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is _____.
15. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
16. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is _____.
17. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
18. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is _____.

19 2 Hydrogen Ions And Acidity Section Review Answers

Linda Swisher, Kevin T. Patton



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