

16.3**COLLIGATIVE PROPERTIES OF SOLUTIONS****Section Review****Objectives**

- Identify the three colligative properties of solutions
- Describe why the vapor pressure, freezing point, and boiling point of a solution differ from those properties of the pure solvent.

Vocabulary

- colligative properties
- freezing-point depression
- boiling-point elevation

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

In a solution, the effects of a nonvolatile 1 on the properties of the solvent are called 2. They include 3 point and vapor pressure 4, and boiling point 5. In each case, the magnitude of the effect is 6 proportional to the number of solute molecules or ions present in the 7.

Colligative properties are a function of the number of solute 8 in solution. For example, one mole of sodium chloride produces 9 as many particles in solution as one mole of sucrose and, thus, will depress the freezing point of water 10 as much.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

11. When added to 1000 g of water, 2 moles of a solute will increase the boiling point by 0.512°C .
12. One mole of solute A will depress the freezing point of 1000 g of water the same as one mole of solute B.

163 Colligative Properties Of Solutions Section Review Answers

Gary D. Hall, Barry S. Reiss



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