

Worksheet 2: Thermodynamics

Entropy and Gibbs Free Energy

- Take a look at the following reactions and predict an entropy change. Will entropy increase (+) or decrease (-).
 - $2\text{NO} (g) + \text{O}_2 (g) \rightarrow 2\text{NO}_2 (g)$
 - $\text{N}_2\text{O} (g) + \text{NO}_2 (g) \rightarrow 3\text{NO} (g)$
 - $\text{H}_2\text{O} (g) \rightarrow \text{H}_2\text{O} (l)$
- The following reactions were performed at 298K. Use the table below to find S° . For the identical reactions to question one, how do the results compare to your predictions?

Standard Molar Entropies $\left(\frac{J}{\text{mol} \cdot K}\right)$	$\text{H}_2\text{O} (g)$: 189	$\text{NO} (g)$: 211	$\text{NO}_2 (g)$: 240	$\text{N}_2\text{O} (g)$: 220
	$\text{H}_2 (g)$: 131	$\text{O}_2 (g)$: 205	$\text{H}_2\text{O} (l)$: 70	$\text{O}_3 (g)$: 239

- $2\text{NO}(g) + \text{O}_2 (g) \rightarrow 2\text{NO}_2 (g)$
 - $\text{N}_2\text{O} (g) + \text{NO}_2 (g) \rightarrow 3\text{NO} (g)$
 - $\text{H}_2\text{O} (g) \rightarrow \text{H}_2\text{O} (l)$
 - $3\text{H}_2 (g) + \text{O}_3 (g) \rightarrow 3\text{H}_2\text{O} (g)$
- What sign must G be for a reaction to be spontaneous?

184 Entropy And Free Energy Worksheet Answers

Lei Huang

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