

**Guided Practice - Acids & Bases****Properties of Acids and Bases**

1. Match the following terms with their descriptions:

- |                        |       |  |
|------------------------|-------|--|
| a. Arrhenius Acid      | _____ | the substance formed when a proton is added to _____       |
| b. Arrhenius Base      | _____ | _____ a base   |
| c. Conjugate Acid      | _____ | the remaining substance when a proton is lost from an acid |
| d. Conjugate Base      | _____ | a substance that produces $\text{OH}^-$ in solution        |
| e. Bronsted-Lowry Acid | _____ | a substance producing $\text{H}^+$ in solution             |
| f. Bronsted-Lowry Base | _____ | a proton acceptor  |
|                        |       | a proton donor   |

2. Which of the following represent conjugate acid-base pairs? Label each species as a conjugate acid or a conjugate base. For those pairs that are  $\text{H}_2\text{O}$  conjugates, write the correct conjugate acid or base for each species in the pair, labeling each species correctly.

- $\text{HSO}_4^-$ ,  $\text{SO}_4^{2-}$
- $\text{HBr}$ ,  $\text{Br}^-$
- $\text{H}_3\text{PO}_4$ ,  $\text{PO}_4^{3-}$
- $\text{H}_2\text{O}$ ,  $\text{HO}^-$

3. Complete the following table regarding acid strength.

Property	Strong Acid	Weak Acid
Definition		
Describe Dissociation		
Species present in solution		
Relative strength of conjugate base compared to water		
Equation showing acid's reaction to water		

# 193 Strengths Of Acids And Bases Answer Key

**Glenn H. Miller**



### 193 Strengths Of Acids And Bases Answer Key:

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