TYPES OF CHEMICAL REACTIONS

DISPLACEMENT (SINGLE AND DOUBLE)

INSTRUCTIONS:

Balance the equations below. Indicate the type of displacement reaction by writing SINGLE or DOUBLE in the space provided. Examples are provided below.

1Cu + 2 AgNO₃ → 2 Ag + 1 Cu(NO₃)₂

Reaction: SINGLE DISPLACEMENT

1 CuSO₄ + 2 NaOH → 1 Na₂SO₄ + 1 Cu(OH)₂

Reaction: DOUBLE DISPLACEMENT

QUESTIONS:

1)
$$_Cl_2 + _KI \rightarrow _KCI + _I_2$$
 _____DISPLACEMENT

2)
$$_K_2CO_3 + _BaCl_2 \rightarrow _KCl + _BaCO_3$$
 _____DISPLACEMENT

3)
$$Na + MgCl_2 \rightarrow NaCl + Mg$$
 _____DISPLACEMENT

5)
$$_Al + _Pb(NO_3)_2 \rightarrow _Al(NO_3)_3 + _Pb$$
 ____DISPLACEMENT

6)
$$_AgNO_3 + _CaBr_2 \rightarrow _AgBr + _Ca(NO_3)_2$$
 _____DISPLACEMENT

7)
$$_Al_2(SO_4)_3 + _Ca_3(PO_4)_2 \rightarrow _AlPO_4 + _CaSO_4$$
 $___DISPLACEMENT$

8)
$$_FeBr_2 + _K_2CO_3 \rightarrow _FeCO_3 + _KBr$$
 _____DISPLACEMENT

10) __CuS + __KCl
$$\rightarrow$$
 __CuCl₂ + __K₂S _____DISPLACEMENT

Double Displacement Reactions Answer Key

Disha Experts

Double Displacement Reactions Answer Key:

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