

INORGANIC CHEMISTRY 201 COURSE NOTE (Pearson International Edition)

Chapter 1 Elements- Periodic Table arrangement, Physical Properties

Matter consist of - Elements (simplest form of matter)

There are 117 elements known for now e.g. H, Na, Cl, O₂, N, C etc.

All known elements are arranged in increasing atomic number in the Periodic Table

All elements are represented by a symbol (abbreviation that is used to represent element)

Symbols consist of one or two letters – H, Na, N, Ca

Element: are made up of atoms

Atoms are the smallest particle or basic unit of an element that can take part in a chemical reaction

- Has same characteristics as the element
- Each element has its own characteristic atom different from those of other elements
- Atom is from the word ATOMOS – minute particles
- Atoms are tiny particles not visible to the naked eyes but can be viewed by powerful microscope such as Scanning Tunneling Microscope (STM)
- For example, 1.0 g of Na and 16 g of Oxygen contain 2.6×10^{23} atoms and 6.02×10^{23} atoms respectively
- Then the mass of one atom of Na is 3.8175×10^{-23} and that of oxygen 2.66×10^{-23}
- These measurements were made with highly precise instruments such as a mass spectrometer and nuclear reaction energies

Atomic Mass Scale

- The mass of an atom is in the range of 1.0×10^{-25} . This is so small that it may not be useful or convenient in normal laboratory scale
- Physical scientist and chemist devised a relative scale which is more suitable for calculating atomic masses. This scale is called the **atomic mass scale (ams)**
- Atomic masses are measured relative to a reference element that is arbitrarily chosen for convenience sake and assigned an arbitrary number.

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Terry C. Jones



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Table of Contents Download Course Outline For Futa Chemistry 201 Course

1. Understanding the eBook Download Course Outline For Futa Chemistry 201 Course
 - The Rise of Digital Reading Download Course Outline For Futa Chemistry 201 Course
 - Advantages of eBooks Over Traditional Books
2. Identifying Download Course Outline For Futa Chemistry 201 Course
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Download Course Outline For Futa Chemistry 201 Course
 - User-Friendly Interface
4. Exploring eBook Recommendations from Download Course Outline For Futa Chemistry 201 Course
 - Personalized Recommendations
 - Download Course Outline For Futa Chemistry 201 Course User Reviews and Ratings
 - Download Course Outline For Futa Chemistry 201 Course and Bestseller Lists

5. Accessing Download Course Outline For Futa Chemistry 201 Course Free and Paid eBooks
 - Download Course Outline For Futa Chemistry 201 Course Public Domain eBooks
 - Download Course Outline For Futa Chemistry 201 Course eBook Subscription Services
 - Download Course Outline For Futa Chemistry 201 Course Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Download Course Outline For Futa Chemistry 201 Course Compatibility with Devices
 - Download Course Outline For Futa Chemistry 201 Course Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Download Course Outline For Futa Chemistry 201 Course
 - Highlighting and Note-Taking Download Course Outline For Futa Chemistry 201 Course
 - Interactive Elements Download Course Outline For Futa Chemistry 201 Course
8. Staying Engaged with Download Course Outline For Futa Chemistry 201 Course
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Download Course Outline For Futa Chemistry 201 Course
9. Balancing eBooks and Physical Books Download Course Outline For Futa Chemistry 201 Course
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Download Course Outline For Futa Chemistry 201 Course
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Download Course Outline For Futa Chemistry 201 Course
 - Setting Reading Goals Download Course Outline For Futa Chemistry 201 Course
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Download Course Outline For Futa Chemistry 201 Course
 - Fact-Checking eBook Content of Download Course Outline For Futa Chemistry 201 Course
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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