

Worksheet – Gas Laws III

Gas Laws: Combined Gas Law

The following problems may be solved using the relationships described by either Charles' or Boyles' laws, together in the combined gas law. Show all the steps to the solution of the problem.

$$V_1P_1/T_1 = V_2P_2/T_2$$

1. Carbon dioxide occupies a 2.54 dm³ container at STP. What will be the volume when the pressure is 150 KPa and 26°C?

$$V_2 = \frac{[2.54 \text{ dm}^3][101.3 \text{ KPa}][26^\circ\text{C} + 273\text{K}]}{[150 \text{ KPa}][273 \text{ K}]} = 1.88 \text{ dm}^3$$

2. Oxygen occupies a fixed container of 5.5L at STP. What will happen to the pressure if the temperature rises to 300K?

$$P_2 = \frac{[5.5 \text{ L}][101.3 \text{ KPa}][300 \text{ K}]}{[5.5 \text{ L}][273 \text{ K}]} = 111.3 \text{ KPa}$$

3. Methane is compressed in a closed 15.8 dm³ container at 101.3 KPa. If the volume drops to 8.7 dm³ and the temperature begins at 25°C and then drops to 18°C , what will the pressure of the gas be?

$$P_2 = \frac{[101.3 \text{ KPa}][15.8 \text{ dm}^3][300 \text{ K}]}{[8.7 \text{ dm}^3][273 \text{ K}]} = 179.6 \text{ KPa}$$

4. A helium balloon is fully inflated at 1.2 L. When the clerk is filling the balloon, she stops to make sure it is not going to explode and checks the pressure. The pressure is 5.7 KPa and the temperature at the store is 24°C. The balloon is 0.75L full. She stops when the balloon is fully inflated and the pressure is 7.2 KPa. When the customer takes the balloon outside, it explodes. What was the temperature outdoors?

$$T = \frac{[24^\circ\text{C} + 273\text{K}][1.2 \text{ L}][7.2 \text{ KPa}]}{[0.75 \text{ L}][5.7 \text{ KPa}]} = 600 \text{ K}$$

5. Oxygen gas is added to a rebreather at the same rate as it would be in the open air. The volume is 4.8 dm³ and the temperature is 25 °C. The pressure is 1 atm. Once the tank is lowered under the water the pressure increases to 2 atm and the temperature drops to 12°C. Is the tank in danger of exploding if it's maximum volume is 5.0 dm³?

$$V_2 = \frac{[4.8 \text{ dm}^3][1 \text{ atm}][12^\circ\text{C} + 273\text{K}]}{[2 \text{ atm}][25^\circ\text{C} + 273\text{K}]} = 2.1 \text{ dm}^3, \text{ not in danger of exploding}$$

6. Ozone is formed due to electricity passing through the air and splitting some O₂ atoms which then join briefly to form O₃, ozone. Your olfactory sense can detect 1.00 mL ozone in 10 L of air at STP. You are surprised that the noreaster is accompanied by

Documentary About Gas Laws And Scuba Diving Worksheet Answers

Jicheng Xie



Documentary About Gas Laws And Scuba Diving Worksheet Answers:

The book delves into Documentary About Gas Laws And Scuba Diving Worksheet Answers. Documentary About Gas Laws And Scuba Diving Worksheet Answers is a crucial topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Documentary About Gas Laws And Scuba Diving Worksheet Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Documentary About Gas Laws And Scuba Diving Worksheet Answers
 - Chapter 2: Essential Elements of Documentary About Gas Laws And Scuba Diving Worksheet Answers
 - Chapter 3: Documentary About Gas Laws And Scuba Diving Worksheet Answers in Everyday Life
 - Chapter 4: Documentary About Gas Laws And Scuba Diving Worksheet Answers in Specific Contexts
 - Chapter 5: Conclusion
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- This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Documentary About Gas Laws And Scuba Diving Worksheet Answers.

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