

Chemistry 20L
Practice Problem Set 1
Answer Key
Strong Acids and Bases

1. How many milliliters of 0.0400 M $\text{Ba}(\text{OH})_2$ are required to reach the equivalence point with 15.0 mL of 0.100 HCl?

Answer: Barium hydroxide furnishes two equivalents of hydroxide in solution for each mole of barium hydroxide that you started with, and HCl donates one mole of protons/mole of acid. Thus the concentration of $\text{OH}^- = 0.0800 \text{ N}$. Since equivalents of acid = equivalents of base,

$$(X \text{ mL})0.0800 = (15.0 \text{ mL})(0.100 \text{ N}) \text{ and } X = 18.75 \text{ mL}$$

2. What volume of 0.250 M H_2SO_4 is required to reach the equivalence point with 38.40 mL of 0.200 M NaOH?

Answer: Sulfuric acid furnishes two equivalents of protons for each mole of acid. Thus the concentration of $\text{H}^+ = 0.500 \text{ N}$. Since equivalents of acid = equivalents of base,

$$(X \text{ mL})0.500 \text{ N} = (38.40 \text{ mL})(0.200 \text{ N}) \text{ and } X = 15.4 \text{ mL}$$

3. What volume of 0.120 N H_2SO_4 is required to reach the equivalence point with 24.0 mL of 0.200 N NaOH?

Answer: 40.00 mL (note both concentrations are in normality)

4. How many grams of barium hydroxide are contained in 250 mL of 0.0800 N $\text{Ba}(\text{OH})_2$ solution?

Answer: Note this concentration is normality. $0.0800 \text{ N} = 0.0400 \text{ M}$

Amount of barium hydroxide in 250 mL of 0.0400 M = 0.0100 moles or 1.71 g

4. What weight of NaOH is required to neutralize 250 mL of 0.050 M HCl?

Answer: moles of HCl = moles of NaOH = 0.0125 moles

Weight of NaOH = 0.50 g

6. A sample of an organic acid weighing 0.3446 g was titrated to a phenolphthalein end point with 41.51 mL of 0.1000 N NaOH. Calculate the equivalent weight of the acid.

Answer: equivalents of acid = equivalents of NaOH

$$= (0.04151 \text{ L})(0.1000 \text{ N}) = 0.004151$$

Chemistry Practice Problem Answer Key

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Table of Contents Chemistry Practice Problem Answer Key

1. Understanding the eBook Chemistry Practice Problem Answer Key
 - The Rise of Digital Reading Chemistry Practice Problem Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Practice Problem Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Practice Problem Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Practice Problem Answer Key
 - Personalized Recommendations
 - Chemistry Practice Problem Answer Key User Reviews and Ratings
 - Chemistry Practice Problem Answer Key and Bestseller Lists
5. Accessing Chemistry Practice Problem Answer Key Free and Paid eBooks
 - Chemistry Practice Problem Answer Key Public Domain eBooks
 - Chemistry Practice Problem Answer Key eBook Subscription Services
 - Chemistry Practice Problem Answer Key Budget-Friendly Options
6. Navigating Chemistry Practice Problem Answer Key eBook Formats

- ePub, PDF, MOBI, and More
- Chemistry Practice Problem Answer Key Compatibility with Devices
- Chemistry Practice Problem Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Practice Problem Answer Key
 - Highlighting and Note-Taking Chemistry Practice Problem Answer Key
 - Interactive Elements Chemistry Practice Problem Answer Key
- 8. Staying Engaged with Chemistry Practice Problem Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Practice Problem Answer Key
- 9. Balancing eBooks and Physical Books Chemistry Practice Problem Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Practice Problem Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Practice Problem Answer Key
 - Setting Reading Goals Chemistry Practice Problem Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Practice Problem Answer Key
 - Fact-Checking eBook Content of Chemistry Practice Problem Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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government scholarship enabled rivera to study art at the academy of san carlos in mexico city from age 10 and a grant from the governor of veracruz enabled him to continue his

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