

SS1 Chemistry questions 2nd term

Objectives questions

1. Boyle's law states that at constant temperature, the volume of a gas is: a) Directly proportional to pressure b) Inversely proportional to pressure c) Independent of pressure d) Directly proportional to temperature
2. The equation $P_1V_1 = P_2V_2$ represents: a) Charles' law b) Boyle's law c) Gay-Lussac's law d) Avogadro's law
3. If a gas occupies 4.0 L at a pressure of 2 atm, what will be its volume at 8 atm, assuming constant temperature? a) 1.0 L b) 2.0 L c) 4.0 L d) 8.0 L
4. A sample of gas has a volume of 400 cm³ at 27°C and 1 atm pressure. What will be its volume at 127°C under the same pressure? a) 300 cm³ b) 400 cm³ c) 500 cm³ d) 600 cm³
5. The ideal gas equation is: a) $PV=nRT$ b) PV/T c) $np=RT$ d) $pV=Vp$
6. According to Graham's law, the rate of diffusion of a gas is: a) Directly proportional to its molar mass b) Inversely proportional to the square root of its molar mass c) Directly proportional to the square root of its molar mass d) Inversely proportional to its molar mass
7. If 100 cm³ of hydrogen diffuses through a membrane in 10 seconds, how long will it take for the same volume of oxygen to diffuse? (Molar mass of H₂ = 2, O₂ = 32) a) 20 seconds b) 40 seconds c) 80 seconds d) 160 seconds
8. What is the molar volume of a gas at STP? a) 12.4 dm³ b) 22.4 dm³ c) 32.0 dm³ d) 44.8 dm³
9. Avogadro's law states that equal volumes of gases at the same temperature and pressure contain equal: a) Masses b) Densities c) Number of molecules d) Number of atoms
10. If 2.0 moles of a gas occupy 44.8 dm³ at STP, what is the volume occupied by 5.0 moles at STP? a) 22.4 dm³ b) 44.8 dm³ c) 56.0 dm³ d) 112.0 dm³
11. The most suitable method to separate a mixture of sand and salt is: a) Filtration and evaporation b) Distillation c) Chromatography d) Sublimation
12. The process of heating a liquid to form vapor and condensing the vapor to recover the liquid is: a) Filtration b) Crystallization c) Distillation d) Decantation

Chemistry Questions For Ss1 2nd Term Exam

Kingsley Augustine



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