

# Stoichiometry Worksheet and Key

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1. How many moles of  $\text{O}_2$  will be formed from 1.65 moles of  $\text{KClO}_3$ ?

1.65 mol $\text{KClO}_3$	_____ mol $\text{O}_2$	= _____ mol $\text{O}_2$
	_____ mol $\text{KClO}_3$	

2. How many moles of  $\text{KClO}_3$  are needed to make 3.50 moles of  $\text{KCl}$ ?

3.50 mol $\text{KCl}$		= _____ mol $\text{KClO}_3$

3. How many moles of  $\text{KCl}$  will be formed from 2.73 moles of  $\text{KClO}_3$ ?

_____		= _____



4. How many moles of  $\text{Fe}_2\text{O}_3$  are produced when 0.275 moles of  $\text{Fe}$  is reacted?

0.275 mol $\text{Fe}$		= _____ mol $\text{Fe}_2\text{O}_3$

5. How many moles of  $\text{Fe}_2\text{O}_3$  are produced when 31.0 moles of  $\text{O}_2$  is reacted?

_____		= _____

6. How many moles of  $\text{O}_2$  are needed to react with 8.9 moles of  $\text{Fe}$ ?

_____		= _____



7. How many moles of  $\text{O}_2$  are produced when 1.26 moles of  $\text{H}_2\text{O}$  is reacted?

8. How many moles of  $\text{H}_2\text{O}$  are needed to produce 55.7 moles of  $\text{H}_2$ ?

9. If enough  $\text{H}_2\text{O}$  is reacted to produce 3.40 moles of  $\text{H}_2$ , then how many moles of  $\text{O}_2$  must have been made?  
(a bit challenging, but just think about it and you can probably figure it out)

# Chemistry Stoichiometry Practice Answers

**Morris Hein, Susan Arena, Cary Willard**

## **Chemistry Stoichiometry Practice Answers:**

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*Progress in Ecological Stoichiometry* Dedmer B. Van de Waal, James J. Elser, Adam C. Martiny, Robert W. Sterner, James B. Cotner, 2018 Ecological stoichiometry concerns the way that the elemental composition of organisms shapes their ecology It deals with the balance or imbalance of elemental ratios and how that affects organism growth nutrient cycling and the interactions with the biotic and abiotic worlds The elemental composition of organisms is a set of constraints through which all the Earth s biogeochemical cycles must pass All

organisms consume nutrients and acquire compounds from the environment proportional to their needs Organismal elemental needs are determined in turn by the energy required to live and grow the physical and chemical constraints of their environment and their requirements for relatively large polymeric biomolecules such as RNA DNA lipids and proteins as well as for structural needs including stems bones shells etc These materials together constitute most of the biomass of living organisms Although there may be little variability in elemental ratios of many of these biomolecules changing the proportions of different biomolecules can have important effects on organismal elemental composition Consequently the variation in elemental composition both within and across organisms can be tremendous which has important implications for Earth s biogeochemical cycles It has been over a decade since the publication of Sterner and Elser s book Ecological Stoichiometry 2002 In the intervening years hundreds of papers on stoichiometric topics ranging from evolution and regulation of nutrient content in organisms to the role of stoichiometry in populations communities ecosystems and global biogeochemical dynamics have been published Here we present a collection of contributions from the broad scientific community to highlight recent insights in the field of Ecological Stoichiometry

Foundations of College Chemistry Morris Hein,Susan Arena,Cary Willard,2023 Foundations of College Chemistry 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need This International Adaptation offers new and updated content with improved presentation of all course material It builds on the strengths of previous editions including clear explanations and step by step problem solving The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry There is new and expanded coverage of polarizing power and polarizability Fajans rules collision number and mean free path abnormal molecular masses and van t Hoff factor and applications of radioactivity

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