

# QUIZZZ

AP Chemistry Unit 8 Review  
 8/17 questions · 24  
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 @ 1000000 · 10/2/19

100%   
 100%   
 100%

1.



2.

## B.1 Introduction to Acids & Bases

### Autoionization of Water:



This is an endothermic process—increasing the heat will shift the reaction to produce more  $\text{H}_3\text{O}^+$  and  $\text{OH}^-$ .  $K_w$  increases (more products), both pH and pOH will decrease but your  $\text{H}_3\text{O}^+$  will still be neutral.

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-]$$

Only digits to the right of the decimal are significant in pH and pOH values. See

$$\text{pH} = -\log$$

In case of acidic solution  $> [\text{H}^+] > [\text{OH}^-]$   
 $\text{pH} < 7$  at 25°C  
 In case of neutral solution  $> [\text{H}^+] = [\text{OH}^-]$   
 $\text{pH} = 7$  at 25°C  
 In case of basic solution  $> [\text{H}^+] < [\text{OH}^-]$   
 $\text{pH} > 7$  at 25°C

### ION PRODUCT CONSTANT:

$$K_w = [\text{H}^+][\text{OH}^-] = 1.0 \times 10^{-14} \text{ at } 25^\circ\text{C}$$

$$= K_a \times K_b$$

$$\text{pH} = -\log[\text{H}^+], \text{pOH} = -\log[\text{OH}^-]$$

$$14 = \text{pH} + \text{pOH}$$

Also known:  
 $[\text{H}^+] = 10^{-\text{pH}}$  and  $[\text{OH}^-] = 10^{-\text{pOH}}$

3. The value of  $K_w$  for water at 0°C is  $1 \times 10^{-15}$ . What is the pOH of water at 0°C?

- 7.5  
 7.0

- 8.0  
 6.5

# Chemistry Unit 8 Quiz 1a Answers

**Jessica J Manson**



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