Chp 8 Study guide:

Using a T-chart on a <u>separate piece</u> of binder paper, write the question and then answer the question.

- 1) A chemical equation like Na + Cl -> NaCl proves what law?
- 2) If you have 50 grams of a reactant you will end up with 50 grams of a product. What law is this?
- 3) How does the following equation break the law of conservation of matter?

$$H_2 + O_2 \rightarrow 2H_2O$$

- Use the following equation to answer the questions below CH₄ + 2O₂ → CO₂ + 2H₂O
 - a) How many products produced?
 - b) How many reactants are produced?
- 5) For an exothermic reaction, the heat content of the is greater than that of the Draw a diagram to support this.
- For an endothermic reaction, the heat content of the _____ is greater than that of the ______. Draw a diagram to support this.
- When two are more elements react chemically what can you say about their chemical and physical properties in comparison to the original element.
- 8) How can elements form compounds/molecules? Explain
 - a) Chemical reactions
 - b) Physical changes
 - c) State of changes

- 9) List the types of evidence that will indicate a chemical change.
- 10) If one of your reactants weighs 50 grams and your product weighs 100 grams, what is the mass of your other reactants? What law is this?
- 11) Give three examples of a chemical change.
- 12) Give three examples of a physical change.
- 13) Give two examples of an endothermic reaction.
- 14) Give two examples of an exothermic reaction.
- 15) What type of change has occurred in the below equation, a chemical or physical change? Explain

Balance the following equations

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Michael Rajnik

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