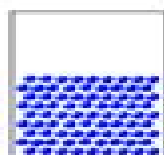


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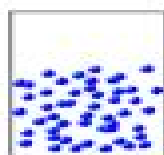
Kinetic Molecular Theory Worksheet

Use the following terms to fill in the blanks. Some may be used more than once. Hopefully all will be included but you never know. If you seem to need an additional term (oops!) just do your best

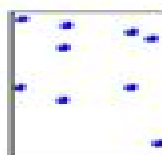
Gas	Solid	Liquid	Phase	Compress
Empty Space	Kinetic Energy	Bonds	Structure	Collisions
Anomaly	Random Motion	Temperature	Absolute Zero	density
Fluids	Compressible	Incompressible		



A



B



C

A	SOLID
B	LIQUID
C	GAS

The boxes above represent matter in the 3 common phases. Write the name of the phase next to the appropriate letter.

- The word kinetic comes from a Greek word that means "to move." The kinetic molecular theory is based upon the assumption that atoms are in constant **Random Motion**.
- State the names of the phases in order of Temperature, Kinetic Energy, density and bond strength between atoms

Lowest Temperatures	In between Temperatures	Highest Temperatures
Solid	Liquid	Gas
Lowest Kinetic Energy	In between Kinetic Energy	Highest Kinetic Energy
Solid	Liquid	Gas
Lowest Density		Highest Density
Gas		
High Strength Bonds	Weak Bonds	No Bonds
Solid	Liquid	Gas

- Kinetic Energy** is directly related to temperature. As the temperature increases so does the **Kinetic Energy**.
- Compared to Solids and Liquids gases tend to have a lower **density**. This can be explained by the fact that there is so much **Empty Space** between particles in a gas.
- A small pressure will easily **compress** a gas. If this happens its **density** will increase. The space between the atoms will (INCREASE or **DECREASE**).
- Even a large pressure will not **compress** a solid because the atoms are so closely packed.
- Solids have strong **Bonds**, liquids have weaker **Bonds** and gases have no **Bonds** holding them together.
- A solid like ice or diamond has a highly ordered arrangement of bonds between atoms which is called its **Structure**. A liquid or gas does not have this.

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