

Kinetics, Equilibrium and Thermodynamics: Practice Test

1. Multiple Choice

1. An enzyme is a biological catalyst. Which of the following is true?

- It provides an alternative mechanism.
- It provides an alternative mechanism.
- It provides a pathway of activation that reduces the activation energy of a reaction.
- It provides the energy for the reaction to proceed.
- It is not specific.

2. A catalyst increases the rate of a reaction by:

- increasing the activation energy.
- decreasing the activation energy.
- increasing the activation energy of the reaction.
- decreasing the activation energy of the reaction.
- none of these.

3. A catalyst increases the rate of a reaction by:

$$\text{Step 1: } \text{A} + \text{B} \rightarrow \text{C} + \text{D} \quad \Delta H = -100 \text{ kJ mol}^{-1}$$

$$\text{Step 2: } \text{C} + \text{D} \rightarrow \text{E} + \text{F} \quad \Delta H = -100 \text{ kJ mol}^{-1}$$

$$\text{Overall: } \text{A} + \text{B} \rightarrow \text{E} + \text{F} \quad \Delta H = -200 \text{ kJ mol}^{-1}$$

4. The overall reaction is exothermic. The activation energy is:

- 100 kJ mol⁻¹
- 200 kJ mol⁻¹
- 100 kJ mol⁻¹
- 200 kJ mol⁻¹
- 100 kJ mol⁻¹

5. A catalyst increases the rate of a reaction by:

$$\text{Step 1: } \text{A} + \text{B} \rightarrow \text{C} + \text{D} \quad \Delta H = -100 \text{ kJ mol}^{-1}$$

$$\text{Step 2: } \text{C} + \text{D} \rightarrow \text{E} + \text{F} \quad \Delta H = -100 \text{ kJ mol}^{-1}$$

$$\text{Overall: } \text{A} + \text{B} \rightarrow \text{E} + \text{F} \quad \Delta H = -200 \text{ kJ mol}^{-1}$$

6. The overall reaction is exothermic. The activation energy is:

- 100 kJ mol⁻¹
- 200 kJ mol⁻¹
- 100 kJ mol⁻¹
- 200 kJ mol⁻¹
- 100 kJ mol⁻¹

7. The overall reaction is exothermic. The activation energy is:

- 100 kJ mol⁻¹
- 200 kJ mol⁻¹
- 100 kJ mol⁻¹
- 200 kJ mol⁻¹
- 100 kJ mol⁻¹



8. The overall reaction is exothermic. The activation energy is:

$$\text{Step 1: } \text{A} + \text{B} \rightarrow \text{C} + \text{D} \quad \Delta H = -100 \text{ kJ mol}^{-1}$$

$$\text{Step 2: } \text{C} + \text{D} \rightarrow \text{E} + \text{F} \quad \Delta H = -100 \text{ kJ mol}^{-1}$$

9. The overall reaction is exothermic. The activation energy is:

- The temperature is high enough to provide enough energy to overcome the activation energy.
- The temperature is high enough that the activation energy is overcome.
- The temperature is high enough to provide enough energy to overcome the activation energy.
- The temperature is high enough to provide enough energy to overcome the activation energy.

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phenomena Accessible and easy to read throughout Surface Science provides a solid foundation from which to develop a conceptual understanding of the field Divided into four sections Surface Science begins with descriptions of the structure thermodynamics and mobility of clean surfaces then moves on to explore the interaction of gas molecules with solid surfaces Next Professor Hudson discusses the energetic particle interactions that are the basis for the majority of techniques developed to reveal the structure and chemistry of surfaces The book concludes with a presentation of the background material involved in crystal nucleation and growth The product of more than three decades of experience in introducing students to surface science this book includes State of the art surface analysis techniques Examples of phenomena and structures from current and classical works A comprehensive presentation that can be easily tailored to senior undergraduate and graduate courses in a variety of disciplines Extensive references and course proven end of chapter problems Surface Science is an excellent textbook for advanced undergraduate and graduate students in engineering and the physical sciences who want a general overview of surface science It also provides important background information for researchers just starting out in the field JOHN B HUDSON PhD is Professor of Materials Science in the Department of Materials Science and Engineering at Rensselaer Polytechnic Institute Troy New York *CliffsAP Chemistry, 4th Edition*

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