Thermochemistry Test Preview

Matching

		Match each item with the correct statement below		
		a. calorimeter d. b. calorie e.		
		c. joule f.	heat capacity	
		c. jouic	near capacity	
 quantity of heat needed to raise the temperature of 1 g of water by 1°C 		1 g of water by 1°C		
	2.	SI unit of energy		
	3.			
	 quantity of heat needed to change the temperature of an object by 1°C 			
device used to measure the heat absorbed or released during a chemical or physical p			sed during a chemical or physical process	
	6.	 heat content of a system at constant pressure 		
		Match each item with the correct statement below		
		a. heat of reaction d.		
		b. heat of formation e.		
		e. Hess's law of heat summation		
	7.	the enthalpy change for a chemical reaction exact	v as it is written	
	8.	이 보고 있는데 있다. 전투에 경기를 가면하고 있다면서 그렇게 하고 있는데 얼마에게 같아야 한다면 하다면 하다면 하다면 하다면 하다면 하다.		
	the energy required to melt a solid at its melting point			
the change in enthalpy that accompanies the formation of a compound from its elements				
	11.	 states that if you add two or more thermochemical equations to give a final equation, you can also ade beats of reaction to give the final heat of reaction 		
		Thoice letter of the choice that best completes the stateme	nt or answers the question.	
	1.75	What happens to the energy produced by burning	Persiana in a sua analona	
	T. day	a. The energy is lost as heat in the exhaust.	gasonite in a car engine:	
		b. The energy is transformed into work to move	the car.	
		c. The energy heats the parts of the engine.		
		d. all of the above		
	1.3.	A piece of metal is heated, then submerged in coo	I water. Which statement below describes what happens?	
		a. The temperature of the metal will increase.	그리다 그리다 그리다 하는 것이 되고 있다면 그 것이 되었다. 그리다 하는 그리다 하는 것이 없는 것이 없다.	
		 The temperature of the water will increase. 		
		 The temperature of the water will decrease. 		
		d. The temperature of the water will increase and	I the temperature of the metal will decrease.	
	1.4.	How does a calorie compare to a joule?		
		 a. A calorie is smaller than a joule. c. 	A calorie is equal to a joule.	
		 b. A calorie is larger than a joule. d. 	The relationship cannot be determined.	
	1.5.			
		The flask would probably feel cooler than before the reaction started.		
		 The flask would probably feel warmer than before the reaction started. 		
		 The flask would feel the same as before the re 	action started.	
		d. none of the above		

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