Equilibrium Practice Problems

1. Write the equilibrium expression for each of the following reactions:

$$N_{2 (g)} + 3 H_{2 (g)} \leftrightarrow 2 NH_{3 (g)}$$

$$K = \frac{[NH_3]^2}{[N_2][H_2]^3}$$

$$I_{2(x)} + CI_{2(g)} \leftrightarrow 2ICI_{(g)}$$

$$K = \underline{\prod_{[Cl_2]}}$$

$$NO_{2(g)} \leftrightarrow NO_{(g)} + 1/2 O_{2(g)}$$

$$K = [NO][O_2]^{1/2}$$

 $[NO_2]$

 The dissociation of acetic acid, CH₃COOH, has an equilibrium constant at 25°C of 1.8 x 10⁻⁵. The reaction is

If the equilibrium concentration of CH_3COOH is 0.46 moles in 0.500 L of water and that of CH_3COO' is 8.1 x 10^{-3} moles in the same 0.500 L, calculate [H $^+$] for the reaction.

$$K = [CH_3COO^*][H^*]$$

$$[CH_3COOH]$$

Plug in known values and solve.

$$1.8 \times 10^{-5} = [8.1 \times 10^{-3} \text{ moles/0.500 L}] [H^*]$$

[0.46 moles/0.500 L]

$$[H^+] = 1.0 \times 10^{-3} M$$

Indicate the effect of a catalyst, pressure, temperature and concentration on each of the following on:

		catalyst	pressure	temperature	concentration
a.	speed	presence increase	direct relationship	direct relationship	direct relationship
ь.	pos"	nothing	based on moles	based on enthalpy of rxn	inverse relationship

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Table of Contents Kp Equilibrium Practice Problems With Answers

- 1. Understanding the eBook Kp Equilibrium Practice Problems With Answers
 - The Rise of Digital Reading Kp Equilibrium Practice Problems With Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Kp Equilibrium Practice Problems With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Kp Equilibrium Practice Problems With Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Kp Equilibrium Practice Problems With Answers
 - Personalized Recommendations
 - Kp Equilibrium Practice Problems With Answers User Reviews and Ratings
 - Kp Equilibrium Practice Problems With Answers and Bestseller Lists

- 5. Accessing Kp Equilibrium Practice Problems With Answers Free and Paid eBooks
 - Kp Equilibrium Practice Problems With Answers Public Domain eBooks
 - Kp Equilibrium Practice Problems With Answers eBook Subscription Services
 - Kp Equilibrium Practice Problems With Answers Budget-Friendly Options
- 6. Navigating Kp Equilibrium Practice Problems With Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Kp Equilibrium Practice Problems With Answers Compatibility with Devices
 - Kp Equilibrium Practice Problems With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Kp Equilibrium Practice Problems With Answers
 - Highlighting and Note-Taking Kp Equilibrium Practice Problems With Answers
 - Interactive Elements Kp Equilibrium Practice Problems With Answers
- 8. Staying Engaged with Kp Equilibrium Practice Problems With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Kp Equilibrium Practice Problems With Answers
- 9. Balancing eBooks and Physical Books Kp Equilibrium Practice Problems With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Kp Equilibrium Practice Problems With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Kp Equilibrium Practice Problems With Answers
 - Setting Reading Goals Kp Equilibrium Practice Problems With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Kp Equilibrium Practice Problems With Answers
 - Fact-Checking eBook Content of Kp Equilibrium Practice Problems With Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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